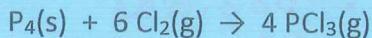
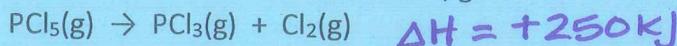
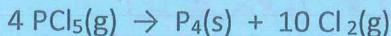


Name BETH "KEY" Period _____
 Hess's Law WS

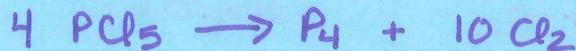
- (1) Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



$$\Delta H = -2439 \text{ kJ}$$



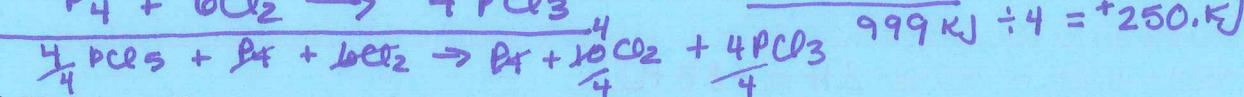
$$\Delta H = 3438 \text{ kJ}$$



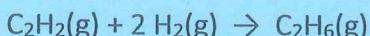
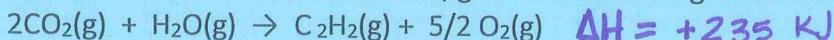
$$3438 \text{ kJ}$$



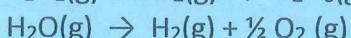
$$-2439 \text{ kJ}$$



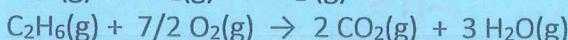
- (2) Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



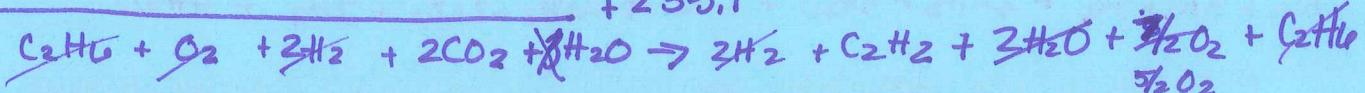
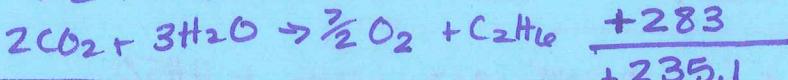
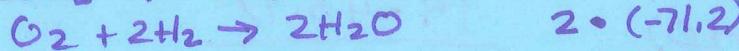
$$\Delta H = -94.5 \text{ kJ}$$



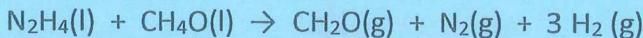
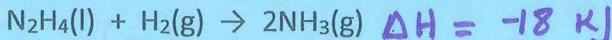
$$\Delta H = 71.2 \text{ kJ}$$



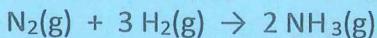
$$\Delta H = -283 \text{ kJ}$$



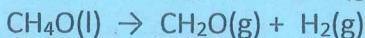
- (3) Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



$$\Delta H = -37 \text{ kJ}$$



$$\Delta H = -46 \text{ kJ}$$



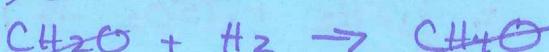
$$\Delta H = -65 \text{ kJ}$$



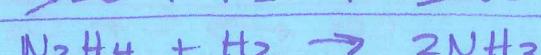
$$-37$$



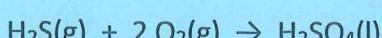
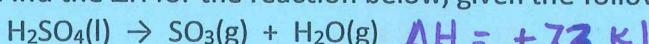
$$-46$$



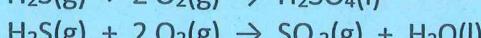
$$+65$$



- (4) Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



$$\Delta H = -235.5 \text{ kJ}$$



$$\Delta H = -207 \text{ kJ}$$



$$\Delta H = 44 \text{ kJ}$$



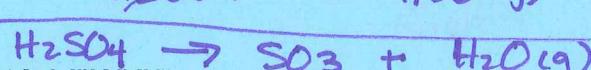
$$+235.5 \text{ kJ}$$



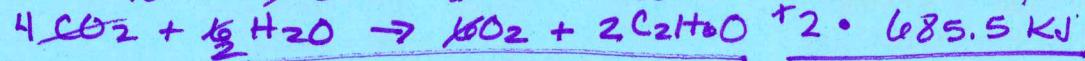
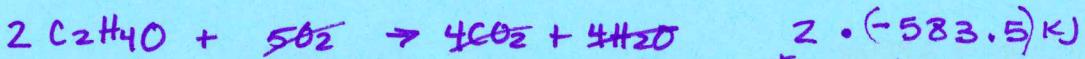
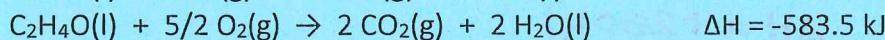
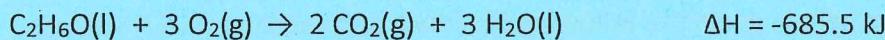
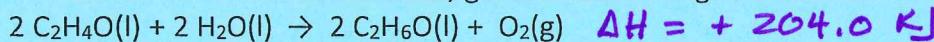
$$-207 \text{ kJ}$$



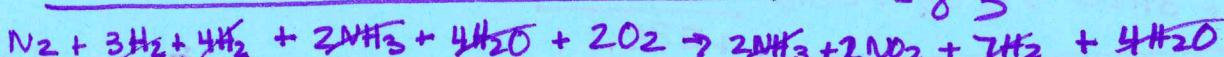
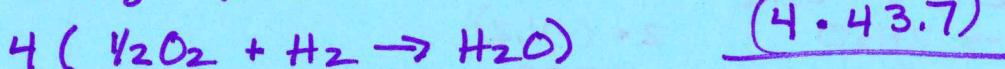
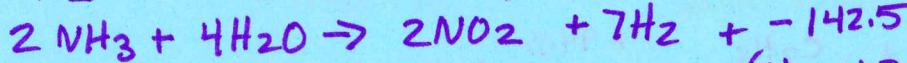
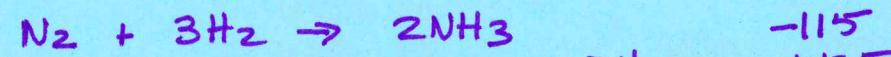
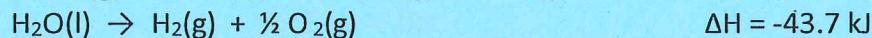
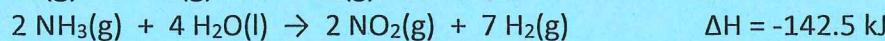
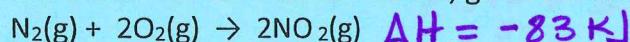
$$44 \text{ kJ}$$



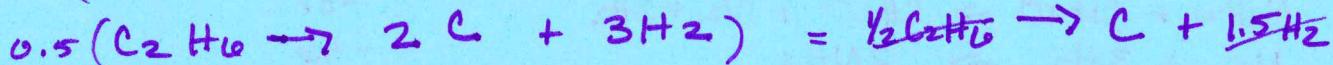
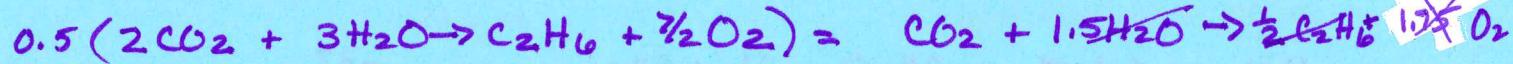
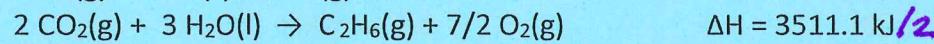
- (5) Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



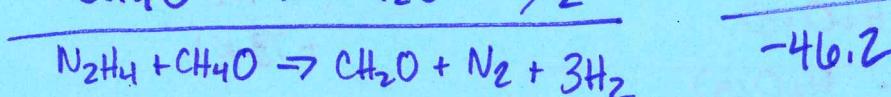
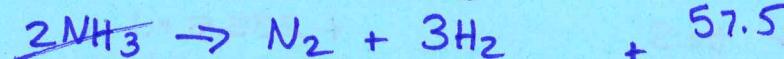
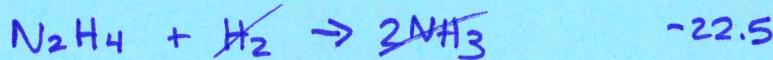
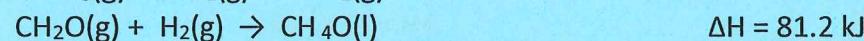
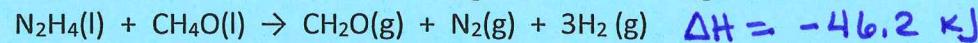
- (6) Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



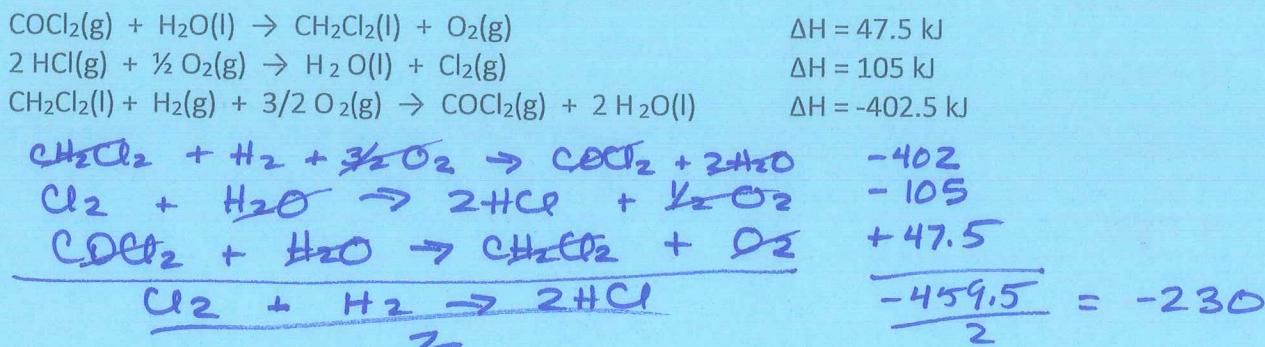
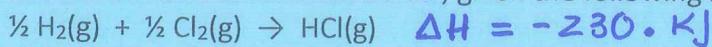
- (7) Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



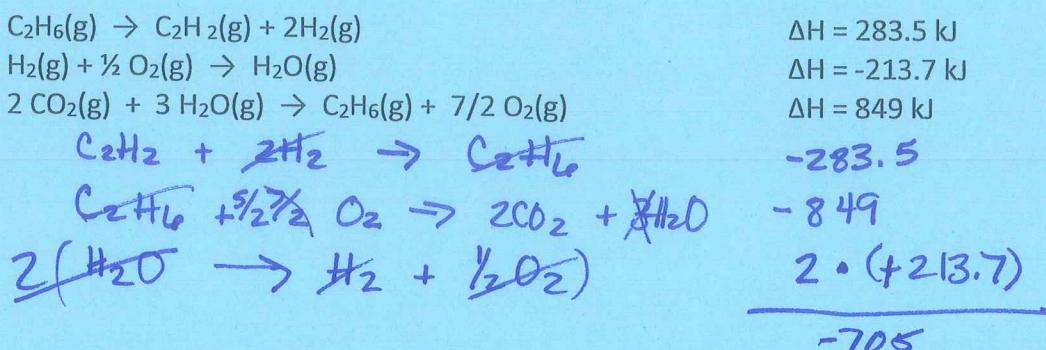
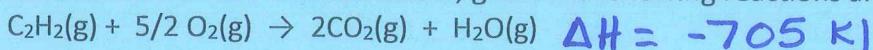
- (8) Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



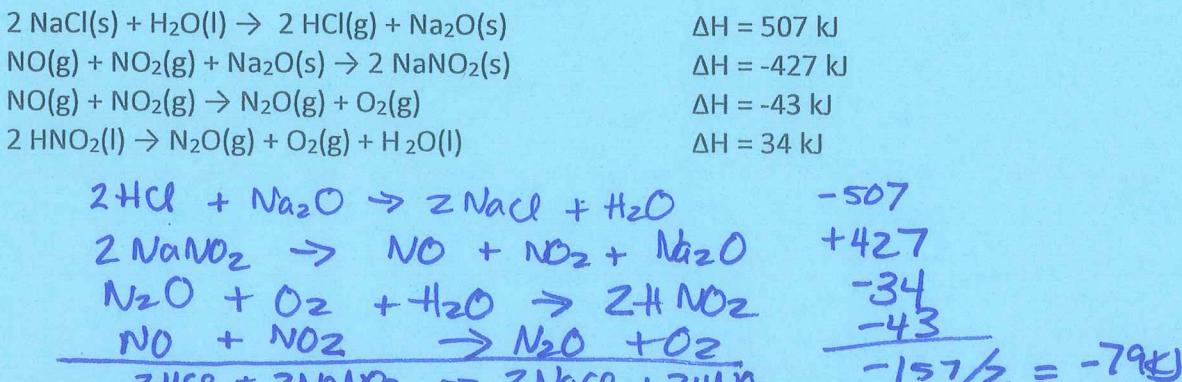
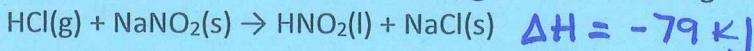
- (9) Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



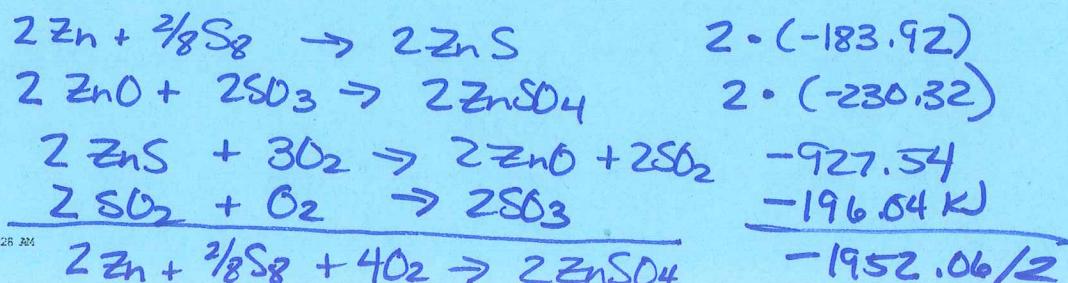
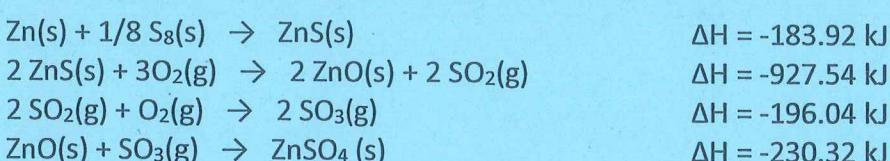
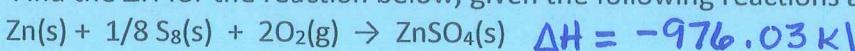
- (10) Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



- (11) Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



- (12) Find the ΔH for the reaction below, given the following reactions and subsequent ΔH values:



Answers:

1. +250. kJ

2. +235 kJ

3. -18 kJ

4. +73 kJ

5. +204.0 kJ

6. -83 kJ

7. +886 kJ

8. -46.2 kJ

9. -230. kJ

10. -705 kJ

11. -78 kJ

12. -976.03 kJ